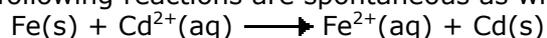
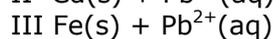
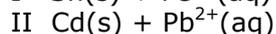
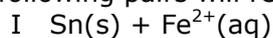


### Exercise 8.32 – Reaction feasibility

**Q832-01** The following reactions are spontaneous as written



Which of the following pairs will react spontaneously?



- A. I only  
B. II only  
C. III only  
D. II and III only

**Q832-02** Use these equations which refer to aqueous solutions to answer the questions which follow:



Deduce whether a silver coin will react with aqueous magnesium chloride solution. [2]

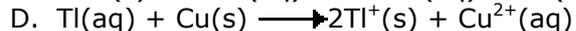
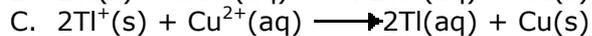
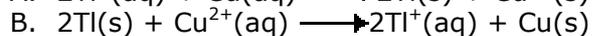
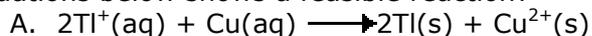
**Q832-03** A part of the reactivity series of metals in order of decreasing reactivity is shown below:

- Magnesium
- Zinc
- Iron
- Lead
- Copper
- Silver

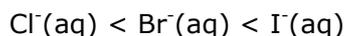
If a piece of copper metals were placed in separate solution of zinc nitrate and silver nitrate:

- a) Determine which solution would undergo reaction? [1]  
b) Identify the type of chemical change taking place in the copper and write a half equation for this reaction. [2]  
c) State, giving a reason, what visible change would take place in the solution. [2]

**Q832-04** Given the information that thallium is a more reactive metal than copper, which of the equations below shows a feasible reaction?



**Q832-05** Investigation tells us that the order of reducing strength of the halide ions (lower to higher) is in the order:



Which of the following shows a feasible reaction for halogens with halide ions:

