

Exercise 8.13 – Oxidation and reduction reactions

Q813-01 Which species is reduced in the following reaction:



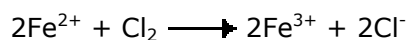
Q813-02 Which species is oxidised in the following reaction:



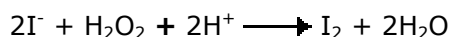
Q813-03 Identify the species being oxidised in the following equation:



Q813-04 Identify the species being reduced in the following ionic equation:



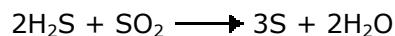
Q813-05 Identify the species being oxidised in the following ionic equation:



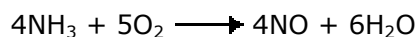
Q813-06 Identify the reducing agent in the following equation:



Q813-07 Identify the oxidising agent in the following equation:



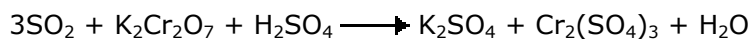
Q813-08 Identify the reducing agent in the following equation:



Q813-09 Identify the oxidising agent in the following equation:



Q813-10 Identify the reducing agent in the following equation:



Extension Write out the half-equations for the oxidation and reduction process happening in each reaction (do not split up covalent substances). Identify any spectator ions.
