

Exercise 6.23 – The position of equilibrium

Q623-01 For a reaction which goes to completion the equilibrium constant, K_c , is:

- A. $\gg 1$
 - B. $\ll 1$
 - C. $= 1$
 - D. $= 0$
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Q623-02 Which statement is true about chemical reactions at equilibrium?

- A. The forward and backward reactions proceed at equal rates
 - B. The forward and backward reactions have stopped
 - C. The concentrations of the reactants and products are equal
 - D. The forward reaction is exothermic
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Q623-03 The position of a reversible reaction initially at equilibrium is shifted to the right until it reaches a new equilibrium situation. Which statement must be true for the reaction when the new position of equilibrium is reached?

- A. The rate of the forward reaction is greater than the rate of the reverse reaction.
 - B. The concentration of reactants and products do not change
 - C. The concentrations of reactants and products are equal
 - D. The value of K_c is greater than 1
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