Exercise 3.24 – Electronegativity

Q324-01 Which series is ranked in order of increasing electronegativity?

A. O, S, Se, Te
B. Cl, S, P, Si
C. In, Sn, N, O
D. C, Si, P, Se

Q324-02 For which pair of atoms is the electronegativity difference the greatest?

A. B, C
B. Li, I
C. K, Cl
D. Se, S

Q324-03 The correct ordering of the electronegativities of the following atoms is:

A. N < O < H
B. H < N < O
C. H < O < N
D. O < N < H

Q324-04 Which statement about electronegativity is correct?

A. Electronegativity decreases across a period
B. Electronegativity increases down a group
C. Metals generally have lower electronegativity values than non-metals
D. Noble gases have the highest electronegativity values

Q324-05 As atomic number increases within a group the electronegativity of the elements

A. Decreases because the atomic number increases
B. Decreases because the atomic size increases
C. Increases because the number of energy levels increases
D. Increases because the atomic number increases
