

Exercise 3.11– The periodic table

Q311-01 Which element should have properties most like those of phosphorus?

- A. Si
 - B. S
 - C. At
 - D. Sb
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Q311-02 Atoms of elements in a group on the Periodic Table have similar chemical properties. This similarity is most closely related to the atoms'

- A. number of principal energy levels
 - B. number of valence electrons
 - C. atomic numbers
 - D. atomic masses
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Q311-03 In which pair are the elements most similar in their chemical properties?

- A. B and N
 - B. Li and Fr
 - C. Mg and Al
 - D. S and Cl
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Q311-04 For which element are the group number and the period number the same?

- A. Li
 - B. Be
 - C. B
 - D. Mg
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Q311-05 Which is related to the number of electrons in the outer main energy level of the elements from the alkali metals to the halogens?

I - group number
II - period number

- A. I only
 - B. II only
 - C. Both I and II
 - D. Neither I nor II
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Q311-06 An element E, of mass number 40 has the electronic configuration 2,8,8,2. Which statement regarding this element is not correct

- A. It belongs to group 2 of the periodic table
 - B. It has 20 neutrons
 - C. It belongs to period 4 of the periodic table
 - D. The formula of its oxide is EO₂
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Q311-07 An element is in group 3 and period 2. How many electrons are present in its outer shell?

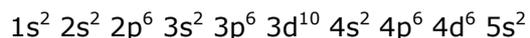
- A. 2
 - B. 3
 - C. 5
 - D. 6
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Q311-08 In which group of the Periodic Table are you most likely to find a metalloid?

- A. The alkali metal family
 - B. The alkaline earth family
 - C. The carbon family
 - D. The halogen family
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Q311-09 In which region of the periodic table would the element with the electronic structure shown below be located?



- A. group 6
 - B. noble gases
 - C. 's' block
 - D. 'd' block
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Q311-10 Which one of the following properties of elements is best used as a basis for classifying them into groups on the periodic table?

- A. Electronic configurations
 - B. Mass number
 - C. Relative atomic mass
 - D. Atomic sizes
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