

Exercise 1.37 – Equations from description

Q137-01 Zinc metal reacts with copper sulphate solution to produce solid copper metal and zinc sulphate solution.

Q137-02 Solid calcium hydroxide reacts with solid ammonium chloride on heating to produce solid calcium chloride, steam and ammonia gas.

Q137-03 When magnesium nitrate is heated in a dry tube magnesium oxide, nitrogen dioxide gas and oxygen are produced.

Q137-04 Liquid silicon tetrachloride reacts with water to produce silicon dioxide and hydrogen chloride gas.

Q137-05 When a solution of calcium hydrogen carbonate is heated a precipitate of calcium carbonate is produced together with carbon dioxide gas and water.

Q137-06 When propane (C_3H_8) gas is burnt with excess air, carbon dioxide and water vapour are produced.

Q137-07 Chlorine gas reacts with potassium iodide solution to give a brown solution of iodine and potassium chloride solution.

Q137-08 Sodium metal reacts with water to produce a solution of the hydroxide and hydrogen gas.

Q137-09 Combustion of methane gas gives carbon dioxide gas and water vapour.

Q137-10 Solid manganese (IV) oxide oxidises hydrochloric acid producing chlorine gas, water and manganese (II) chloride solution.
