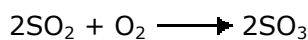


Exercise 1.35 – Particles in equations

Q135-01 How many molecules of ammonia are produced according to the following reaction?



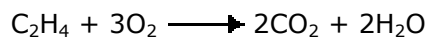
Q135-02 How many molecules of sulphur dioxide are reacting according to the following equation?



Q135-03 How many molecules of oxygen are reacting in the following equation?



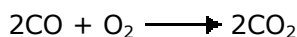
Q135-04 How many molecules of carbon dioxide are formed in the following reaction?



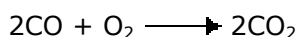
Q135-05 In the reaction between ammonia and oxygen shown below, how many oxygen molecules would be needed to react completely with 8 ammonia molecules?



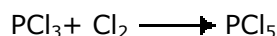
Q135-06 In the following equation how many moles of oxygen are needed to completely react with 2 moles of carbon monoxide?



Q135-07 In the following equation how many moles of carbon dioxide are formed when excess oxygen completely reacts with 10 moles of carbon monoxide?

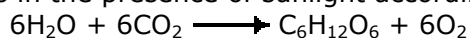


Q135-08 State how many moles of chlorine react completely with 0.1 moles of phosphorus trichloride according to the following equation:



Q135-09 How many moles of oxygen are required to completely react with 5 moles of ethene C_2H_4 if the only products are carbon dioxide and water?

Q135-10 In a photosynthesis reaction glucose molecules are made by combining carbon dioxide molecules and water molecules in the presence of sunlight according to the equation:



How many molecules of oxygen are formed when 20 molecules of carbon dioxide reacts with excess water?
