

## Exercise 0.52 – Metallic bonding

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**Q052-01** What are responsible for the high electrical conductivity of metals?

- A. Delocalised positive ions
  - B. Delocalised valence electrons
  - C. Delocalised atoms
  - D. Delocalised negative ions
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**Q052-02** Which is a correct description of metallic bonding?

- A. Positively charged metal ions are attracted to negatively charged ions
  - B. Negatively charged metal ions are attracted to positively charged metal ions
  - C. Positively charged metal ions are attracted to delocalised electrons
  - D. Negatively charged metal ions are attracted to delocalised electrons
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